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24 ChemQuest 29 Monc LcW s Stnuctunes tr l Name Date:  
Hour: Introduction Questions Draw the Lewis structures for  
the following molecules. Part c is done for you. a) H<sub>2</sub>O  
(oxygen is central) b) SO<sub>2</sub> c) N<sub>2</sub> It t-t (a [Jrt @ N=N @ @ f / @  
2 In question 1 above you should see three different kinds  
of bonds of bonds as best you can. a) single bond: one es.iv  
(ee-) shave a b) \*-l; double bond:

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ChemQuest 29 Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_

Introduction Questions 1. Write the equation for the reaction

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of sodium and chlorine (diatomic) to form sodium chloride.  
2. Write the equation for the reaction of calcium nitride and sodium chloride to produce calcium chloride and sodium nitride.

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subscript is a small number that tells you how many atoms  
are in a compound. For example, in  $\text{CaCl}_2$  the two is the  
subscript and it tells us that there are two chloride ions  
bonded to one calcium. A coefficient tells also tells us how  
many atoms or compounds there are, but in a different way.

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14. Based on your answer to question 13, verify that this equation is true: This equation is called the “ Combined Gas Law ” . Notice that it contains all of the equations combined into one! Again, the pressure times the volume divided by the Kelvin temperature equals 1.18 for every trial in the table. 15.

## [Information: Gas Pressure](#)

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